

NOMENCLATURE

BINARY

2 kinds of elements
count as one element
OH⁻ hydroxide
CN⁻ cyanide
NH₄⁺ ammonium

FIRST ELEMENT METAL

name of first / stem of second + ide
element element

METAL MORE THAN SINGLE

OX #

name of first / (ox # in rom. num.) / stem of second + ide
element element

IF METAL IS Fe, Cu, Sn, Pb 2 NAMES

**name of ion / stem of second element + ide*

FIRST ELEMENT HYDROGEN

Hydro + stem of second element + ic / acid

NONBINARY

more than 2 kinds of elements

FIRST ELEMENT H

*** MEMORIZE

FIRST ELEMENT NOT H

SECOND ELEMENT NOT H

name of first / ic = ate
element ous = ite

METAL MORE THAN

SINGLE OX # SAME WAY

name metal / (OX #) / ic = ate
ous = ite

Fe, Cu, Sn, Pb TWO NAMES

**name ion / ic = ate*

ous = ite

**** MnO₄⁻, SCN⁻, C₂H₃O₂⁻, C₂O₄²⁻, CrO₄²⁻, Cr₂O₇²⁻

name first element (metals same) / name ion

SECOND ELEMENT IS H

name first / hydrogen or / ic = ate
element dihydrogen ous = ite

METAL MORE A SINGLE OX # SAME WAY

SO_x OR CO_x ALSO NAMED

Name of first element (metals same) / bi + ic = ate
ous = ite

* Cu⁺ cuprous
 Cu⁺² cupric
 Fe⁺² ferrous
 Fe⁺³ ferric
 Sn⁺² stannous
 Sn⁺⁴ stannic
 Pb⁺² plumbous
 Pb⁺⁴ plumbic

****C₂O₄⁻² oxalate
 C₂H₃O₂⁻ acetate
 SCN⁻ thiocyanate
 CrO₄⁻² chromate
 Cr₂O₇⁻² dichromate
 MnO₄⁻ permanganate

**

subscript	prefix
1	Omit (mono)
2	di
3	tri
4	tetra
5	penta
6	hexa
7	hepta
8	octa
9	nona
10	deca

***Outies the ic acid has 3 oxygens

Innies the ic acid has 4 oxygens

One more oxygen than the ic = per + stem of second element + ic/acid

the ic acid = stem of second element + ic/acid

One less oxygen than the ic = stem of second element + ous/acid

Two less oxygen than the ic = hypo + stem of second element + ous/acid